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Szekrényes, Zsolt, Nagy, Péter R., Tarczay, György, Maggini, Laura, Bonifazi, Davide ORCID: <https://orcid.org/0000-0001-5717-0121> and Kamarás, Katalin 2018. Direction-dependent secondary bonds and their stepwise melting in a uracil-based molecular crystal studied by infrared spectroscopy and theoretical modeling. Chemical Physics Letters 691 , pp. 163-168. 10.1016/j.cplett.2017.11.006 file

Publishers page: <http://dx.doi.org/10.1016/j.cplett.2017.11.006>  
<<http://dx.doi.org/10.1016/j.cplett.2017.11.006>>

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# Direction-dependent secondary bonds and their stepwise melting in a uracil-based molecular crystal studied by infrared spectroscopy and theoretical modeling

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## Abstract

Three different types of supramolecular interactions can be identified in crystals of 1,4-bis[(1-hexylurac-6-yl)ethynyl]benzene, a uracil-based molecule with a linear backbone: intermolecular H-bonds along the molecular axis, van der Waals interactions between the hexyl chains connecting these linear assemblies, and  $\pi$  stacking perpendicular to the molecular planes. On heating, two disjoint transitions are observed, disordering of hexyl chains at 473 K, followed by H-bond melting at 534 K. The nature of the bonds and transitions was established by matrix-isolation and temperature-dependent infrared spectroscopy and supported by theoretical computations.

**Keywords:** H-bond, DFT computations, self-assembly

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## 1. Introduction

H-bonds formed in uracil-based molecules have played a major role not only in nucleic acid research [1, 2, 3, 4], but recently also as the secondary forces holding together two-dimensional networks with long-range order [5, 6, 7]. Numerous studies on the role of H-bonding in uracil and uracil derivative molecules have been reported [5, 8, 9, 10, 11, 12, 13, 14, 15]. The study of uracil derivatives is of special interest as these compounds are excellent models to demonstrate the key role of supramolecular chemistry both in biochemistry and materials science, such as the formation of two-dimensional porous networks on surfaces [5, 10, 13, 16, 17, 18, 19, 20].

We present a study on a uracil-based molecular solid, formed of 1,4-bis[(1-hexylurac-6-yl)ethynyl]benzene (2), investigating the nature of the interaction between molecular units in the crystalline environment by infrared spectroscopy and theoretical methods. We have performed this study as a continuation of our work on a simple monouracil derivative (1-hexyl-6-ethynyluracil (1H6EU, 1) [21]. Two 1H6EU units (Fig. 1 (a)), connected by a benzene ring, form 2 (Fig. 1 (b)). Molecule 1 is a monotopic molecule with H-bond-forming sites at one end, a prospective ingredient of dimer structures, while 2 is a ditopic module enabling the formation of a linear backbone as H-bond-forming functional groups are present at its opposite ends. This di-topic uracil derivative unit is a very good candidate for a linear linker in self-assembled supramolecular networks. 2 has already proved its potential in forming 2D bicomponent hexagonal porous structures at the solid-liquid interface [16] as well as in molecular recognition through triple H-bonds [10].

In derivative 1, we studied the type of H-bonds by comparing the matrix-isolation infrared spectra of individual molecules with those of the molecular solid; in addition, we followed the melting process of the H-bonds at elevated temperature by both experimental and theoretical methods [21]. Here we report a similar investigation with the result that just below the H-bond melting temperature, a different phase transition occurs in the solid, from a planar tetrameric to a stacked columnar structure. These two consecutive transitions indicate the relative importance of stacking interactions compared to H-bond/van der Waals forces.

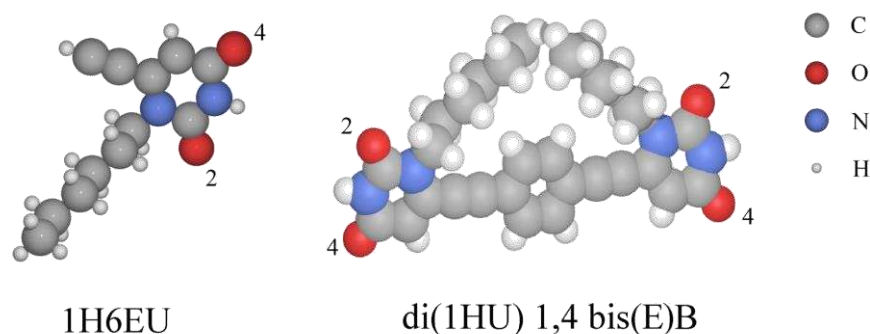


Figure 1: 1 (a) and 2 (b) have identical hydrogen bond forming functional groups (the carbonyl and amine groups). The two carbonyl groups numbered 2 and 4, respectively, may participate in different H-bonding motifs.

## 2. Materials and methods

### 2.1. Materials and experimental methods

Molecules 1 and 2 (Fig. 1) were prepared according to previously published procedures [10, 17]. Temperature-dependent infrared spectra in the solid state were recorded on powders ground in KBr pellets. Mixing with KBr dilutes the material so it becomes transparent to IR light, but the grain size in the pellets ( $\sim 1 \text{ m}^3$ ) ensures that they still can be regarded as solids, preserving the structure.

The matrix isolation (MI) setup is described in detail elsewhere [22, 23]. Briefly, the evaporated sample was mixed with argon (Messer, 99.9997%) before deposition onto an 8-10 K CsI window. The gas flow was kept at  $0.07 \text{ mmol min}^{-1}$ , while the evaporation temperature was optimized to 353.5 K and 530.5 K for molecules 1 and 2, respectively. Under these conditions, the sample consists predominantly of isolated molecules, with only a small amount of aggregated species present. Therefore, comparing MI-IR spectra with solid-state KBr pellet spectra allows us to study the effect of aggregation.

MI-IR spectra were recorded with a Bruker IFS 55 spectrometer with  $1 \text{ cm}^{-1}$  resolution and a DTGS detector while the temperature dependence in KBr pellets with a Bruker IFS66v vacuum FTIR instrument with  $2 \text{ cm}^{-1}$  resolution and an MCT detector. All spectra were taken in the  $400\{4000 \text{ cm}^{-1}$

range with a Ge/KBr beamsplitter. The baseline was corrected by an adjusted polynomial function.

## 2.2. Theoretical methods

Stabilization energies of noncovalent complexes (i.e. dimerization, trimerization, etc. energies) were computed by density functional theory (DFT). Equilibrium structures were optimized using the !B97X-D functional developed by Chai and Head-Gordon [24, 25] and the 6-311G(d,p) basis set. The !B97X-D is a long-range corrected hybrid density functional with dispersion corrections and found to be reliable to describe various noncovalent interactions of molecules of main group elements [26, 27, 28, 29].

For each optimized structure, an additional single-point energy computation was performed with the !B97X-D/6-311++G(3df,3pd) method. All reported stabilization energies are obtained with this protocol, except for the case of the van der Waals tetramer structure. This tetramer illustrates the structure of two neighboring unit cells in our hypothetical solid state structure introduced in more detail in Section 3. However, without the surrounding crystal field, this tetramer geometry does not correspond to a local minimum with respect to the nuclear coordinates. To obtain an approximate substructure characteristic of the three-dimensional lattice, a constrained geometry optimization was performed. The tetramer structure was kept approximately in its plane by applying harmonic constraints on the torsion angles that set the planes of the eight uracil rings parallel to this plane. The obtained structure cannot be considered as a local minimum and accordingly, we do not deduce quantitative information from this result.

Let us note that since the present systems contain several noncovalent interactions of different nature, it is important to take the basis set superposition error (BSSE) into consideration. The BSSE can be reduced significantly using a close to complete basis set, such as 6-311++G(3df,3pd). Intermolecular BSSEs were estimated for dimers of 1 and 2 by computing counterpoise corrections [30], were found to be smaller than 1 kcal mol<sup>-1</sup> in all cases and consequently have been neglected. All computations were performed by the Gaussian 09 program package [31].

In the previous study on 1 systems, temperature dependent ab initio molecular dynamics (AIMD) computations were performed in the canonical ensemble [21]. These results were in good agreement with the experimental spectra and provided important aid to the assignment and understanding of

temperature dependent IR experiments on 1. The large size of the 2 dimers-tetramers prevents us from computing free energies and infrared spectra for these systems, but due to structural similarities, considerations formulated for 1 based on AIMD data may be transferable to 2. For instance, based on the same arguments as presented in Ref. [21], only diketo tautomers were considered in case of all 2 based systems.

### 3. Results and discussion

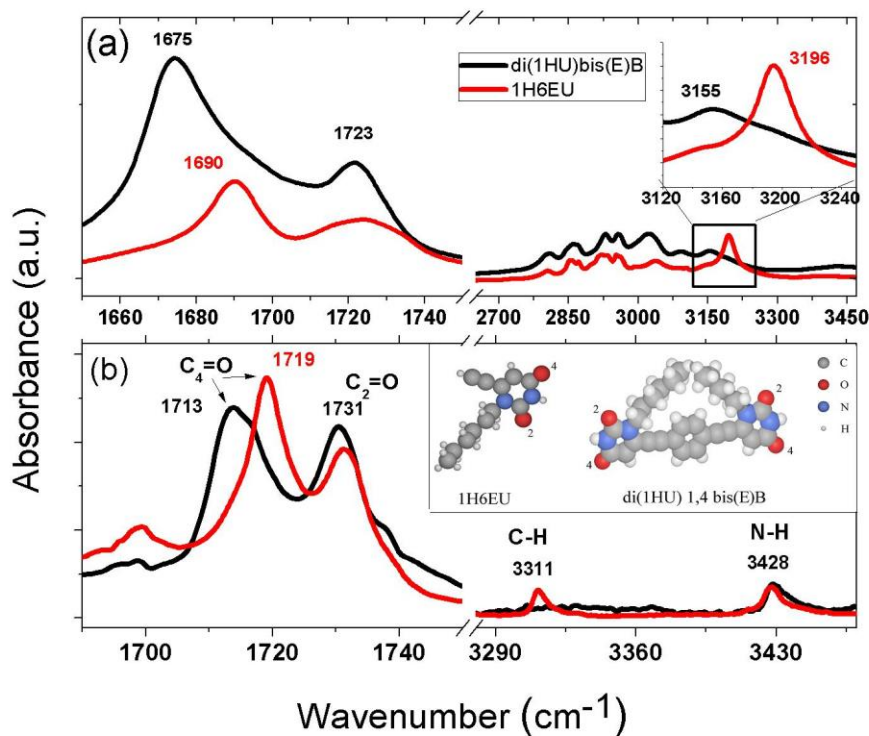


Figure 2: (a) Experimental room-temperature spectra of 1 and 2 and (b) MI-IR spectra of the isolated molecular units. Subscripts 2 and 4 refer to the corresponding carbonyl groups according to Fig. 1. The spectrum for 1 was taken from Ref. [21].

Figure 2 presents the IR spectra of 1 and 2 recorded in the solid state at room temperature and matrix-isolation infrared spectra of isolated monomers

recorded in an argon matrix at 8 K. We previously assigned the IR modes of **1** considering the solid sample as a static assembly of dimers with H-bonds constituting the main intermolecular interaction [21]. Evaluating the possible dimer conformers of **1**, we found that the N-H (3) and one of the C=O groups (2 or 4) is always involved in H-bond formation, while the other C=O group is free. The band at  $1690\text{ cm}^{-1}$  was assigned to the H-bonded carbonyl group while the band centered at  $1723\text{ cm}^{-1}$  represents the free C=O stretching vibration. The band at  $3196\text{ cm}^{-1}$  corresponds to the H-bonded N-H stretch mode. In analogy, for derivative **2** we assign the H-bonded C=O and N-H stretch modes to the bands at  $1675\text{ cm}^{-1}$  and  $3155\text{ cm}^{-1}$  and the free C=O to that at  $1723\text{ cm}^{-1}$ . While the free carbonyl mode is at the same frequency for molecules **1** and **2**, there is a considerable difference for the H-bond affected IR bands. These bands shift to lower wavenumbers for **2** suggesting a stronger interaction of the molecular modules through H-bonds. In the MI-IR spectra (Fig. 2 (b)) the mode assignment of the isolated monomers of derivative **1** was done by comparing with the theoretical spectrum computed for monomers at 5 K. The bands at  $1719\text{ cm}^{-1}$  and at  $1731\text{ cm}^{-1}$  belong to the free C4=O and C2=O (according to Fig. 1) vibrations of the monomer, respectively. For molecule **2** the free vibration of C4=O and C2=O is assigned to the band at  $1713\text{ cm}^{-1}$  and at  $1731\text{ cm}^{-1}$ . The  $6\text{ cm}^{-1}$  difference between the C4=O vibrational frequencies of the two molecules could be attributed to their structural differences. The band situated at  $3428\text{ cm}^{-1}$  is assigned to the free N-H vibration and is at the same frequency for both units. Finally, the band at  $3311\text{ cm}^{-1}$ , present only in **1**, corresponds to the ethynyl stretch mode (C-H).

In Table 1 the calculated dimerization energies for **1** and **2** for three different configurations are presented (according to Fig. 3). Looking at **1** first, the order of stability and the dimerization energies obtained by the !B97X-D/6-311++G(3df,3pd) method match previous theoretical results and experimental observations for **1**. Comparing **1** and **2** we observe the same tendency, the most stable dimer configuration being the (4-4) conformer followed by the (2-4) and (2-2) conformers. The dimerization energy of **2** indicates a slightly stronger coupling through H-bonds than for **1**, in agreement with the IR results presented in Fig. 2.

Figure 4 presents the temperature dependence of the infrared spectrum of **2** from room temperature up to 573 K. There is an intermediate state between the stable H-bonded phase at room temperature and the case of melted H-bonds above 543 K. Between 473 K and 543 K the intense C=O H-bonded



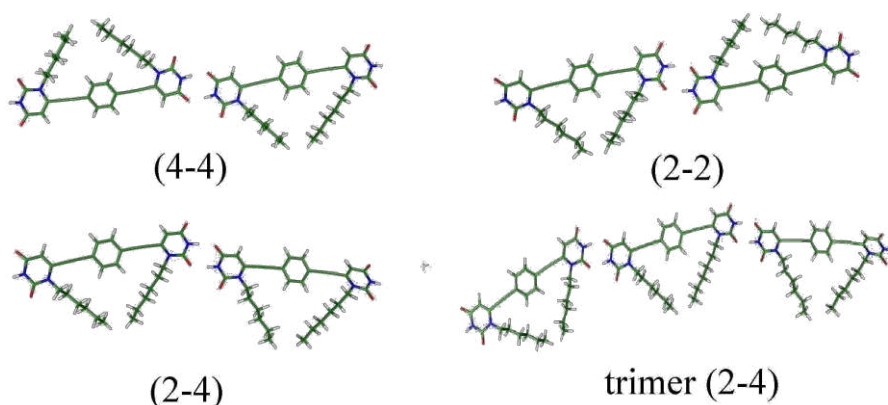


Figure 3: The investigated dimer and trimer configurations of 2 in the theoretical study. For the dimer notation we indicate the labels of the carbon atoms (2 or 4) of the carbonyl groups taking part in the H-bond (according to Fig. 1).

Table 1: Formation energies of the different isolated dimers for 1 and 2.

Dimer	1 (kcal mol <sup>-1</sup> )	2 (kcal mol <sup>-1</sup> )
(4-4)	-13.3	-14.8
(2-2)	-12.4	-13.4
(2-4)	-12.6	-13.9

band (at 1675 cm<sup>-1</sup>) disappears and a new band appears at 1695 cm<sup>-1</sup>. This temperature induced transition can also be characterized by studying the behavior of the amine group. N-H groups are always involved in H-bonding between molecules in this system. One-step melting of hydrogen bonds would result in the appearance of free N-H vibration bands; here, however, the H-bonded N-H band (at 3150 cm<sup>-1</sup>) persists up to 543 K. All this indicates an intermediate phase, i.e. a phase transition at 473 K into a different, but still H-bonded, structure. The H-bonded N-H vibration, similarly to the H-bonded C=O stretch, shifts towards higher wavenumbers at 473 K. Above 543 K the H-bonded N-H band decreases in intensity and a new band appears at 3411 cm<sup>-1</sup>, in the region characteristic of the free N-H vibrations.

In the following, we present a possible three dimensional arrangement of molecules 2 following the experimental and theoretical data. We start from the structure organized on a surface, observed in a previous scanning tunnel-



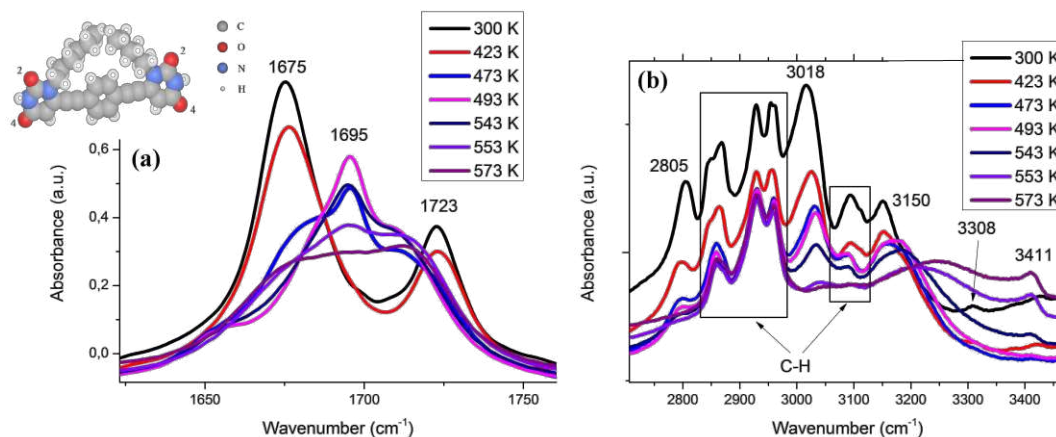


Figure 4: Experimental temperature-dependent IR spectra of 2. Above 473 K the band at  $1675\text{ cm}^{-1}$  disappears and a new band appears at  $1695\text{ cm}^{-1}$  which starts to decrease in intensity above 543 K. In the high frequency region the amine vibrations ( $3150\text{ cm}^{-1}$ ) persist up to 543 K, above this temperature the free amine vibration band appears at  $3411\text{ cm}^{-1}$  indicating the total melting of the H-bonds.

ing microscopy (STM) study [10]. This STM investigation, performed under ultrahigh vacuum on Ag(111) surfaces, provides important insight into the system of noncovalent interactions that governs the self-assembly of molecule 2. The observed double-row supramolecular structure is illustrated in Fig. 5(a) for four units. Although the (4-4) dimer is found to be slightly more pre-ferred energetically (cf. Table 1), the STM image veries that a close-packed crystal structure can only be built with (2-4) type H-bonds. Furthermore, the all-cis con guration of the hexyl chains was found to be more abundant in Ref. [10] even at submonolayer coverage in two dimensions. Therefore we continue with the theoretical study of the supramolecular structure shown in (Fig. 5(a)).

First of all, parallel linear assemblies are created by double hydrogen bonds between molecules 2 (cf. Fig. 3(d)). Theoretical investigation reveals that double H-bonds can be formed independently on both ends of molecule 2. Therefore this linear structure can grow to any length, since the stabilization energy, corresponding to the energy required for the extension of the chain by a single unit, is independent of chain size. This tendency is illustrated for the cases of trimers and tetramers in Table 2.

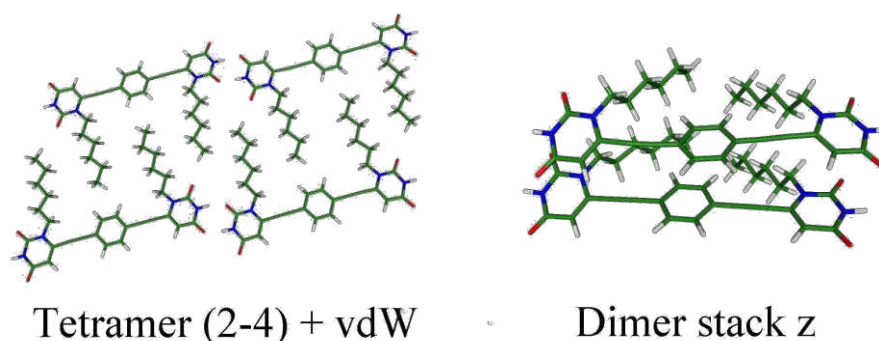


Figure 5: Double-row wire structure organized by double H-bonds and van der Waals interactions between linear molecular assemblies (a). Organization perpendicular to the molecular plane due to  $\pi$  stacking (b).

Table 2: Formation energies for dimer, trimer and tetramer of the (2-4) conformer of 2.

Dimer (kcal mol <sup>-1</sup> )	Trimer (kcal mol <sup>-1</sup> )	Tetramer (kcal mol <sup>-1</sup> )
-12.97	-25.92	-38.85
	( 2 x -12.97)	( 3 x -12.97)

The second organizing principle is the interdigitation of the laterally placed hexyl chains stabilized by van der Waals attractions. To approximate the magnitude of the van der Waals stabilization energy compared to H-bonding, single point DFT computations were carried out on the sub-structures of the tetramer in Fig. 5 (a). Dimerization energies for the van der Waals and H-bonded dimers were found to be -8.3 kcal mol<sup>-1</sup> and - 14.6 kcal mol<sup>-1</sup>, respectively.

Organization in the third direction, perpendicular to the aromatic planes, is governed by  $\pi$ -stacking of the aromatic rings (Fig. 5 (b)). This is supported by previous investigations of  $\pi$ -interactions in uracil based crystals [32] and in stacked uracil dimers [33, 34]. In case of molecule 2,  $\pi$ -stacking is particularly strong due to the three parallel aromatic rings. Dimerization energy for the dimer in Fig. 5 (b) is -33.6 kcal mol<sup>-1</sup>, which consists of mainly  $\pi$ -stacking and partly hexyl-hexyl van der Waals interaction.

Since the van der Waals interaction turns out to be the lowest in energy, it is probable that it provides the "weakest link" to induce the transition at 473 K. Due to the large difference in the stabilization energies, the stronger

double H-bonds are expected to be relatively intact at that temperature. This argument is supported by previous room temperature AIMD IR computations on the (2-4)-type double H-bonded 1 dimer [21]. This structural change can also explain the transition of the H-bonded C=O band from  $1675\text{ cm}^{-1}$  to  $1695\text{ cm}^{-1}$ . At lower temperature the additional van der Waals attraction, being parallel to the double H-bonds, shortens the H-bonds. This is illustrated in table 3 by comparing the computationally obtained geometrical parameters of the (2-4) dimer, the (2-4) linear tetramer and the double-row tetramer of Fig. 5 (a). Structures without the extra van der Waals stabiliza-

tion have H-bond lengths around 1.825 Å and 1.815 Å, while a significant decrease of cca. 0.06 Å can be found for the double-row structure. A shorter H-bond implies a lower force constant and thus lower frequency for the H-bonded C=O vibration, in accord with the experiment.

Table 3: Computed hydrogen bond length values for different complexes of 2.

Hydrogen bond length (Å)	C <sub>4</sub> =(O H)-N	C <sub>2</sub> =(O H)-N
dimer (2-4)	1.825	1.815
tetramer (2-4), linear chain	1.822-1.825	1.816-1.818
tetramer (2-4), double-row	1.76, 1.78	1.73, 1.77

In light of all data, we assign the room temperature spectrum to a crystal structure of ordered double-row linear assemblies. These assemblies are formed via double H-bonds from monomers, and are organized into a double-row structure due to van der Waals interactions between laterally placed hexyl groups. Different layers of this practically planar arrangement are kept together by  $\pi$ -stacking.

#### 4. Conclusion

The supramolecular ordering in the solid state of a bis-uracil based linear molecule has been studied using infrared spectroscopy and theoretical methods. Infrared spectra of the isolated monomers were obtained by the matrix isolation technique. Temperature dependence of the vibrational bands affected by the H-bonds was followed in the solid state. Prior to the total melting of the H-bonds an intermediate state was identified in the temperature profile. By analogy with previous STM measurements [10] and supported by theoretical modeling, we suggest a structural phase transition at 473 K to explain this observation. The low-temperature structure in the

three-dimensional solid is analogous to that on surfaces: double-row, ladder-like arrangements of 2 units bound together by the stacking interactions of the aromatic rings perpendicular to the molecular plane. The two rows of molecules 2 are held together by van der Waals interactions between the hexyl chains. This structure is consistent with temperature dependent infrared and 2D STM measurements and theoretical results as well.

At an intermediate temperature of 473 K the H-bonded C=O band is shifted to a higher wavenumber, but the H-bonds are not disrupted yet, according to the frequency of the amine band. The increase in the C=O frequency can be attributed to the disintegration of the weakest cohesive force, the van der Waals interactions between the hexyl chains. The full melting of the H-bonds was observed at 543 K, where the intensity of the H-bonded amine band decreases and a new band appears in the wavenumber region characteristic of free N-H vibrations. We conclude that the supramolecular order in the three perpendicular directions of the solid crystal is governed by three different noncovalent interactions: double H-bonds, van der Waals attraction and  $\pi$ -stacking. We find this arrangement particularly interesting, since the strength of these interactions varies in a wide range, which may permit direction-specific manipulations of this structure.

## 5. Acknowledgement

This research was funded by the Hungarian Scientific Research Fund (MTA grant no. OTKA K75877, IR measurements: grant No. OTKA NK105697) and supported by the supercomputing facility of the National Information Infrastructure Development Institute (NIIF). The work of PRN is supported through the New National Excellence Program of the Ministry of Human Capacities, ID: NKP-17-4-BME-55. The European Union and the European Social Fund have provided financial support to the project under the grant agreement no. TAMOP 4.2.1./B-09/KMR-2010-0003.

## References

- [1] J. Watson, F. Crick, A structure for deoxyribose nucleic acid, *Nature* 171 (1953) 737{738.
- [2] W. Saenger, *Principles of nucleic acid structure*, Springer, Berlin, 1984.

- [3] G. A. Jeffrey, W. Saenger, Hydrogen bonding in biological structures, Springer-Verlag, Berlin, New York, Heidelberg, 1991.
- [4] G. A. Jeffrey, An introduction to hydrogen bonding, Oxford University Press, Oxford, 1997.
- [5] L. Piot, C.-A. Palma, A. Llanes-Pallas, M. Prato, Z. Szekrenyes, K. Ka-maras, D. Bonifazi, P. Samor , Selective formation of bi-component ar-rays through H-bonding of multivalent molecular modules, *Adv. Funct. Mater.* 19 (2009) 1207{1214.
- [6] J. V. Barth, G. Costantini, K. Kern, Engineering atomic and molecular nanostructures at surfaces, *Nature* 437 (2005) 671{679.
- [7] M. Surin, P. Samor , Multicomponent monolayer architectures at the solid-liquid interface: Towards controlled space-con nected properties and reactivity of functional building blocks, *Small* 3 (2007) 190{194.
- [8] D. Troitino, L. Bailey, F. Peral, Formation of hydrogen-bonded dimers of uracil and amino acids: Cysteine, lysine and phenylalanine. A density functional theory study, *J. Mol. Struc.{Theochem* 767 (2006) 131{138.
- [9] P. Hobza, J. Sponer, Structure, energetics, and dynamics of the nucleic acid base pairs: Nonempirical ab initio calculations, *Chem. Rev.* 99 (1999) 3247{3276.
- [10] A. Llanes-Pallas, M. Matena, T. Jung, M. Prato, M. St•ohr, D. Bonifazi, Trimodular engineering of linear supramolecular miniatures on Ag(111) surfaces controlled by complementary triple hydrogen bonds, *Angew. Chem., Int. Ed.* 47 (2008) 7726{7730.
- [11] B. Feibush, A. Figueroa, R. Charles, K. D. Onan, P. Feibush, B. L. Karger, Chiral separation of heterocyclic drugs by HPLC - solute sta-tionary phase base-pair interactions, *J. Am. Chem. Soc.* 108 (1986) 3310{3318.
- [12] J.-M. Lehn, M. Mascal, A. DeCian, J. Fischer, Molecular recognition directed self-assembly of ordered supramolecular strands by cocrystal-lization of complementary molecular components, *J. Chem. Soc., Chem. Commun.* 1990 (1990) 479{481.

- [13] C.-A. Palma, J. Bjork, M. Bonini, M. S. Dyer, A. Llanes-Pallas, D. Bonifazi, M. Persson, P. Samor , Tailoring bicomponent supramolecular nanoporous networks: The role of phase segregation and polymorphism at the solid-liquid interface, *J. Am. Chem. Soc.* 131 (2009) 13062{13071.
- [14] A. Llanes-Pallas, K. Yoosaf, H. Traboulsi, J. Mohanraj, T. Seldrum, J. Dumont, A. Minoia, R. Lazzaroni, N. Armaroli, D. Bonifazi, Modular engineering of H-bonded supramolecular polymers for reversible functionalization of carbon nanotubes, *J. Am. Chem. Soc.* 133 (2011) 15412{15424.
- [15] L. Dordevic, T. Marangoni, T. Miletic, J. Rubio-Magnieto, J. Mohanraj, H. Amenitsch, D. Pasini, N. Liaros, S. Couris, N. Armaroli, M. Surin, D. Bonifazi, Solvent molding of organic morphologies made of supramolecular chiral polymers, *J. Am. Chem. Soc.* 137 (2015) 8150{ 8160.
- [16] C.-A. Palma, M. Bonini, A. Llanes-Pallas, T. Breiner, M. Prato, D. Bonifazi, P. Samor , Pre-programmed bicomponent porous networks at the solidliquid interface: the low concentration regime, *Chem. Comm.* 2008 (2008) 5289{5291.
- [17] A. Llanes-Pallas, C.-A. Palma, L. Piot, A. Belbakra, A. Lis-torti, M. Prato, P. Samor , N. Armaroli, D. Bonifazi, Engineering of supramolecular H-bonded nanopolygons via self-assembly of programmed molecular modules, *J. Am. Chem. Soc.* 131 (2009) 509{520.
- [18] D. Bonifazi, S. Mohnani, A. Llanes-Pallas, Supramolecular chemistry at interfaces: molecular recognition on nanopatterned porous surfaces, *Chem. Eur. J.* 15 (2009) 7004{7025.
- [19] S. Mohnani, A. Llanes-Pallas, D. Bonifazi, Mastering nanostructured materials through H-bonding recognitions at interfaces, *Pure Appl. Chem.* 82 (2010) 917{929.
- [20] T. Marangoni, D. Bonifazi, Engineering of supramolecular H-bonded nanopolygons via self-assembly of programmed molecular modules, *Nanoscale* 5 (2013) 8837{8851.

- [21] Z. Szekrenyes, K. Kamaras, G. Tarczay, A. Llannes-Pallas, T. Marangoni, M. Prato, D. Bonifazi, J. Bjork, F. Hanke, M. Pers-son, Melting of hydrogen bonds in uracil derivatives probed by infrared spectroscopy and ab initio molecular dynamics, *J. Phys. Chem. B* 116 (2012) 4626{4633.
- [22] G. Pohl, A. Perczel, E. Vass, G. Magyarfalvi, G. Tarczay, A matrix iso-lation study on Ac-Gly-NHMe and Ac-L-Ala-NHMe, the simplest chiral and achiral building blocks of peptides and proteins, *Phys. Chem. Chem. Phys.* 9 (2007) 4698{4708.
- [23] S. Gobi, K. Knapp, E. Vass, Z. Majer, G. Magyarfalvi, M. Hollosi, G. Tarczay, Is beta-homo-proline a pseudo-gamma-turn forming element of beta-peptides? An IR and VCD spectroscopic study on Ac-beta-HPro-NHMe in cryogenic matrices and solutions, *Phys. Chem. Chem. Phys.* 12 (2010) 13603{13615.
- [24] J. D.Chai, M. Head-Gordon, Long-range corrected hybrid density func-tionals with damped atom-atom dispersion corrections, *Phys. Chem. Chem. Phys.* 10 (2008) 6615{6620.
- [25] J. D. Chai, M. Head-Gordon, Systematic optimization of long-range corrected hybrid density functionals, *J. Chem. Phys.* 128 (2008) 084106.
- [26] K. E. Riley, M. Pitonak, P. Jurecka, P. Hobza, Stabilization and struc-ture calculations for noncovalent interactions in extended molecular sys-tems based on wave function and density functional theories, *Chem. Rev.* 110 (2010) 5023{5063.
- [27] L. A. Burns, A. Vazquez-Mayagoitia, B. G. Sumpter, D. C. Sherrill, Density-functional approaches to noncovalent interactions: A compar-ison of dispersion corrections (DFT-D), exchange-hole dipole moment (XDM) theory, and specialized functionals, *J. Chem. Phys.* 134 (2011) 084107.
- [28] L. Goerigk, S. Grimme, A thorough benchmark of density functional methods for general main group thermochemistry, kinetics, and nonco-valent interactions, *Phys. Chem. Chem. Phys.* 13 (2011) 6670{6688.



- [29] O. A. Vydrov, T. Van Voorhis, Benchmark assessment of the accuracy of several van der waals density functionals, *J. Chem. Theory Comput.* 8 (2012) 1929{1934.
- [30] S. Boys, F. Bernardi, The calculation of small molecular interactions by the di erences of separate total energies. some procedures with reduced errors, *Mol. Phys.* 19 (1970) 553{566.
- [31] M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F. Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, J. A. Montgomery, Jr., J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N. Staroverov, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, N. Rega, J. M. Millam, M. Klene, J. E. Knox, J. B. Cross, V. Bakken, C. Adamo, J. Jaramillo, R. Gomperts, R. E. Stratmann, O. Yazyev, A. J. Austin, R. Cammi, C. Pomelli, J. W. Ochterski, R. L. Martin, K. Morokuma, V. G. Za-krzewski, G. A. Voth, P. Salvador, J. J. Dannenberg, S. Dapprich, A. D. Daniels, O. Farkas, J. B. Foresman, J. V. Ortiz, J. Cioslowski, D. J. Fox, Gaussian 09 Revision C.1, gaussian Inc. Wallingford CT 2009.
- [32] Z—. Czyznikowska, R. Zalesny, M. Ziolkowski, R. W. Gora, P. Cysewski, The nature of interactions in uracil dimer: An ab initio study, *Chem. Phys. Lett.* 450 (2007) 132{137.
- [33] M. Pitonak, K. E. Riley, P. Neogady, P. Hobza, Highly accurate CCSD(T) and DFT-SAPT stabilization energies of H-bonded and stacked structures of the uracil dimer, *ChemPhysChem* 9 (2008) 1636{ 1644.
- [34] V. R. Cooper, T. Thonhauser, D. C. Langreth, An application of the van der Waals density functional: Hydrogen bonding and stacking interactions between nucleobases, *J. Chem. Phys.* 128 (2008) 204102.